Cornell Notes	Topic/Objective: **Stoichiometric Conversions**	Name:
AVID Decades *of* College Dreams		Class/Period: **Chemistry**
		Date:

Essential Questions: **How can you predict the amount of reactants needed or products produced?**

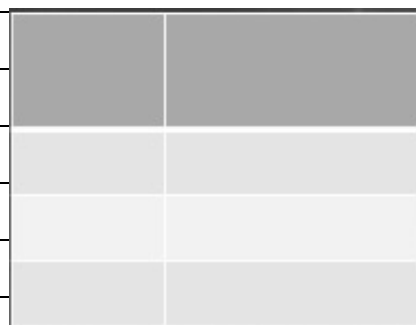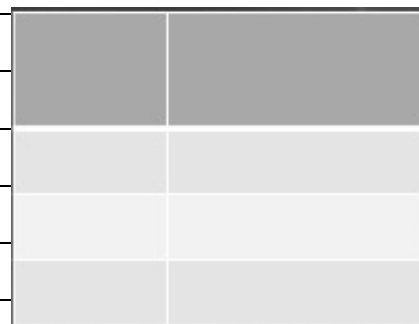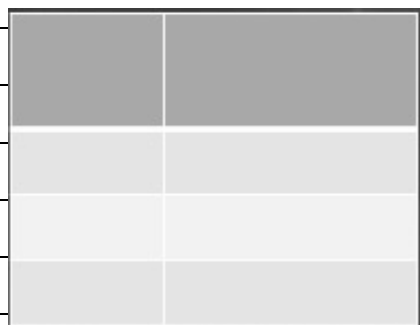| Questions:	Notes:
What is stoichiometry?	
Stoichiometric Mole-Mole Conversion	One disadvantage of burning propane (C_3H_8) is the carbon dioxide (CO_2) is one of the products. The release carbon dioxide increases the growing concentration of CO_2 in the atmosphere. How many moles of carbon dioxide are produced when 10.0 moles of propane are burned in excess oxygen in a gas grill?
Stoichiometric Mole-Mass Conversion	Determine the mass of sodium chloride or table salt (NaCl) produced when 1.25 moles of chlorine gas reacts vigorously with sodium.
11 +1 **Na** Sodium 22.990	
17 +1 +5 +7 -1 **Cl** Chlorine 35.453	

Stoichiometric Mass-Mass Conversion	Ammonium nitrate (NH_4NO_3), an important fertilizer, produces N_2O gas and H_2O when it decomposes. Determine the mass of water produced from the decomposition of 25.0 g of solid ammonium nitrate.

7
N
+1
+2
+3
+4
+5
-1
-2
-3
Nitrogen
14.007 | 2p

1
H
+1
-1
Hydrogen
1.008 | 1s

8
O
-2
Oxygen
15.999 | p

What substitutions can be made in the stoichiometric box?

Summary: