
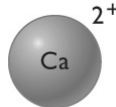
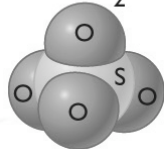


Cornell Notes 	Topic/Objective: <b>Ionic Compound Naming and Formulas with Polyatomic Ions</b>	Name:
		Class/Period: <b>Chemistry</b>
		Date:

Essential Question: **How are ionic compounds (salts) named with polyatomic ions?**

Questions:	Notes:	Monatomic ion 	Polyatomic ion 
<b>What are the different types of ions (not anion/cation)?</b>	<b>Monatomic:</b>		
	<b>Polyatomic:</b>		

<b>How is an ionic compound with a polyatomic ion named?</b>	<ul style="list-style-type: none"> <li>•</li> <li>•</li> <li>○</li> <li>○</li> </ul>
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**Name the following ionic compounds:**



<b>How do you predict the formulas of ionic compounds with polyatomic ions?</b>	

<b>Summary:</b>

# Charges of Common Elements and Polyatomic Ions

1+		2+		3+	
Ammonium	$\text{NH}_4^+$	Barium	$\text{Ba}^{+2}$	Aluminum	$\text{Al}^{+3}$
Copper (I)	$\text{Cu}^+$	Cadmium	$\text{Cd}^{+2}$	Antimony (III)	$\text{Sb}^{+3}$
Gold (I)	$\text{Au}^+$	Calcium	$\text{Ca}^{+2}$	Arsenic (III)	$\text{As}^{+3}$
Hydrogen	$\text{H}^+$	Chromium (II)	$\text{Cr}^{+2}$	Bismuth	$\text{Bi}^{+3}$
Lithium	$\text{Li}^+$	Cobalt (II)	$\text{Co}^{+2}$	Boron	$\text{B}^{+3}$
Potassium	$\text{K}^+$	Copper (II)	$\text{Cu}^{+2}$	Cobalt (III)	$\text{Co}^{+3}$
Sodium	$\text{Na}^+$	Iron (II)	$\text{Fe}^{+2}$	Gold (III)	$\text{Au}^{+3}$
Thalium (I)	$\text{Tl}^+$	Lead (II)	$\text{Pb}^{+2}$	Iron (III)	$\text{Fe}^{+3}$
		Magnesium	$\text{Mg}^{+2}$	Manganese (III)	$\text{Mn}^{+3}$
		Mercury (II)	$\text{Hg}^{+2}$	Thalium (III)	$\text{Tl}^{+3}$
		Nickel (II)	$\text{Ni}^{+2}$	Titanium (III)	$\text{Ti}^{+3}$
		Tin (II)	$\text{Sn}^{+2}$	Zirconium (III)	$\text{Zr}^{+3}$
		Zinc	$\text{Zn}^{+2}$		
1-		2-		3-	
Acetate	$\text{C}_2\text{H}_3\text{O}_2^-$	Carbonate	$\text{CO}_3^{-2}$	Arsenate	$\text{AsO}_4^{-3}$
Bromate	$\text{BrO}_3^-$	Chromate	$\text{CrO}_4^{-2}$	Arsenite	$\text{AsO}_3^{-3}$
Bromide	$\text{Br}^-$	Manganate	$\text{MnO}_4^{-2}$	Borate	$\text{BO}_3^{-3}$
Chlorate	$\text{ClO}_3^-$	Oxalate	$\text{C}_2\text{O}_4^{-2}$	Phosphate	$\text{PO}_4^{-3}$
Chloride	$\text{Cl}^-$	Oxide	$\text{O}^{-2}$		
Cyanide	$\text{CN}^-$	Peroxide	$\text{O}_2^{-2}$		
Hydrogen Carbonate (bicarbonate)	$\text{HCO}_3^-$	Sulfate	$\text{SO}_4^{-2}$		
Hypochlorite	$\text{ClO}^-$	Sulfide	$\text{S}^{-2}$		
Hydroxide	$\text{OH}^-$	Sulfite	$\text{SO}_3^{-2}$		
Nitrate	$\text{NO}_3^-$	Thiosulfate	$\text{S}_2\text{O}_3^{-2}$		
Nitrite	$\text{NO}_2^-$				
Permanganate	$\text{MnO}_4^-$				
Thiocyanate	$\text{SCN}^-$				