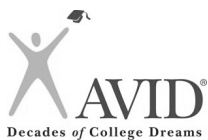


Cornell Notes



Topic/Objective: **Moles and One Step**

Name:

**Problems (Part I)**

Class/Period: **Chemistry**

(Moles, Particles/Moles Conversions, and Molar Mass)

Date:

Essential Questions: **How are things counted in chemistry and their masses determined?**

Questions:

Notes:

**What unit of measure is used in chemistry?**

**Mole:**

**What is a particle?**

**Describe how the conversion chart is used to convert between particles and moles.**

**How do you convert from moles to particles?**

*Calculate the number of molecules in 3.50 moles of sucrose.*

	Formula
Particles	
Particles/Mole	
Moles	

**How do you convert from particles to moles?**

*Zinc is used as a corrosion-resistant coating on iron and steel. It is also an essential trace element in your diet. Calculate the number of moles that contain  $4.50 \times 10^{24}$  atoms of zinc (Zn).*

	Formula
Particles	
Particles/Mole	
Moles	

<p><b>How is the mass of substances calculated?</b></p>	<p><b>Molar Mass:</b></p>
<p><b>How do you determine the mass of an element and a compound?</b></p>	
	<p><b>Examples:</b></p>
	<p>What is the molar mass of lead (Pb)?</p>
	<p>Determine the molar mass of <math>K_2CrO_4</math>.</p>
<p><b>Summary:</b></p>	