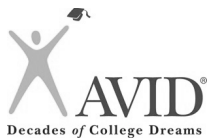


Cornell Notes



Topic/Objective: **Ideal Gas Law**

Name:

Class/Period: **Chemistry**

Date:

Essential Question: **How can you calculate the number of moles of a gas if you know, P, V, and**

T?

Questions:

Notes:

Ideal Gas Law

$n =$

Universal Gas Constant (R) =

Ideal Gas Law Example Problem

How many moles of air are in a 0.5 L breath on top of Mount Everest? The pressure is

0.33 atm and the temperature is 254 K.

Summary: