
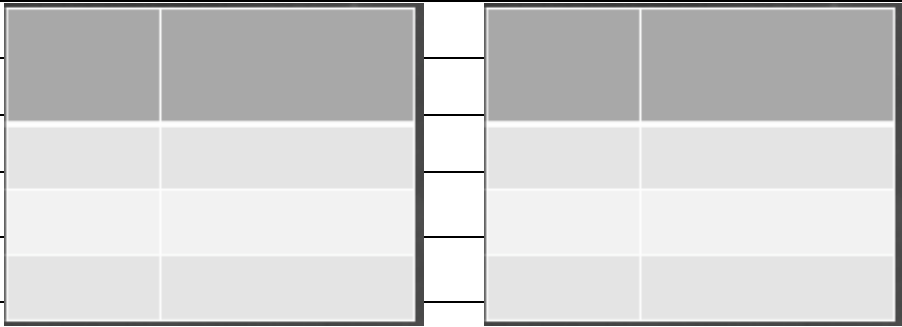


Cornell Notes 	Topic/Objective: The Formula for a Hydrate	Name:
		Class/Period: Honors Chemistry
		Date:

Essential Questions: **How can you determine the formula for a hydrate?**

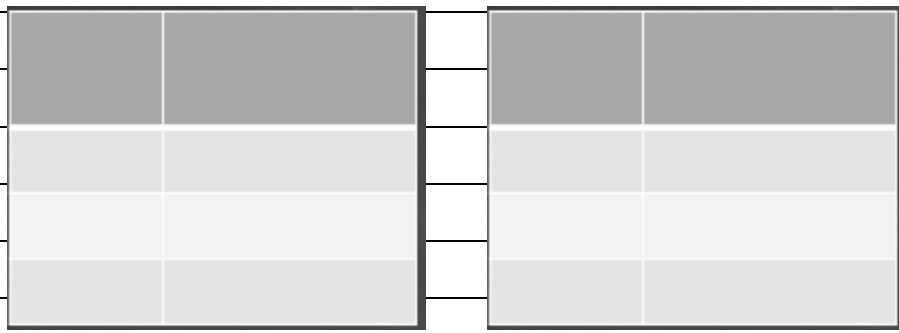
Questions:	Notes:
What is the difference between hydrous and anhydrous?	
How are hydrates named?	
What needs to be done in the lab to determine the formula of a hydrate?	1.
	2.
	
	3.

How do you determine the formula of a hydrate?

12⁺²
Mg
Magnesium
s 24.305

A mass of 2.50 g of powdery, hydrated magnesium sulfate is placed in a crucible and heated. After heating, 1.59 g chunky anhydrous magnesium sulfate ($MgSO_4$) remains. What is the formula for the hydrate? Name the hydrate.

16⁺⁴
⁺⁶
⁻²
S
Sulfur
p 32.066



8⁻²
O
Oxygen
p 15.999

Summary: