Cornell Notes	Topic/Objective: Dalton's Law of Partial Pressures	Name:
		Class/Period: Chemistry
		Date:

Essential Question: How can you calculate the total pressure of a gas from its partial pressures and/or moles?

Questions:	Questions:
What does Dalton's law of partial pressures state?	
	Dalton's Law Equation (partial/total pressures only)
	Dalton's Law Practice Problem
	A mixture of oxygen (O_2), carbon dioxide (CO_2), and nitrogen (N_2) has a total pressure of 0.97 atm. What is the partial pressure of O_2, if the partial pressure of CO_2 is 0.70 atm and the partial pressure of N_2 is 0.12 atm?

Questions:	Notes:	
What equations can be used to calculate partial or total pressures if moles are included in the problem?	First Equation	Second Equation
	X_i =	P_i =
	n_i =	X_i =
	n_T =	P_T =
	Dalton's Law Practice Problem	
	A mixture of gases contains 4.46 moles of neon (Ne), 0.74 mole of argon (Ar), and	
	2.15 moles of xenon (Xe). Calculate the partial pressures of the gases if the total pressure	
	is 2.00 atm at a certain temperature.	

Summary: