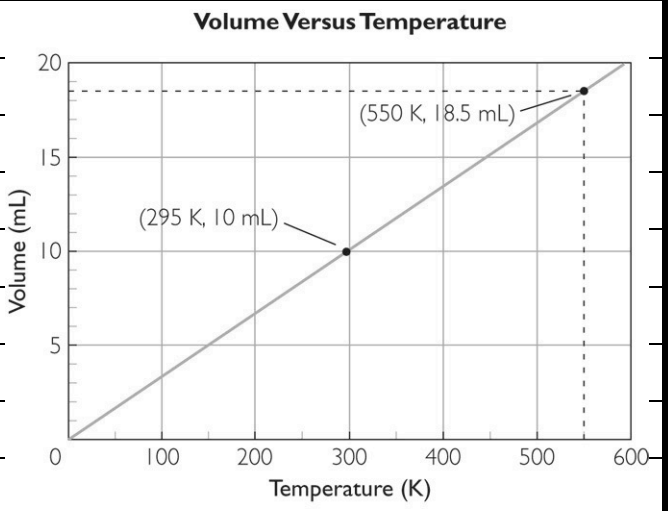
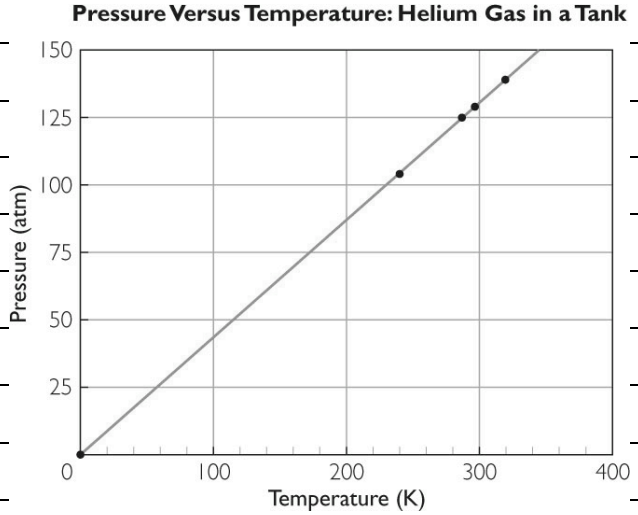
| Cornell Notes | Topic/Objective: Charles' & Gay-Lussac's Law | Name: |
|---|---|---|
| | | Class/Period: Chemistry |
| | | Date: |

Essential Question: How can you predict the volume or pressure of a gas sample?

| Questions: | Notes: |
|---|---|
| What is the relationship between volume and temperature? |   |
| How can volume and temperature be related mathematically? | |
| | **Charles' Law Example Problem** |
| | A gas sample at 40.0° C occupies a volume of 2.32 L. If the temperature is raised to 75.0° C, what will the volume be, assuming the pressure remains constant? |

| Questions: | Notes: |
|---|---|
| What is the relationship between pressure and temperature? |  |
| How can pressure and temperature be related mathematically? | |
| | **Gay-Lussac's Law Example Problem** |
| | A gas sample at -15.0° C has a pressure of 0.25 atm. If the temperature is raised to 155.0° C, what will the new temperature be, assuming the volume remains constant? |



| Summary: |
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