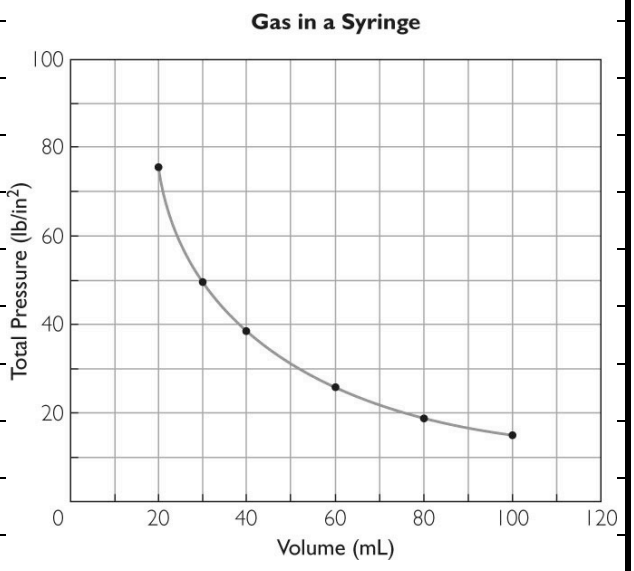
Cornell Notes	Topic/Objective: Boyle's Law	Name:
		Class/Period: Chemistry
		Date:

Essential Question: What is the relationship between pressure and volume and how can that relationship be used mathematically?

Questions:	Notes:
What happens on the molecular level to cause an increase in pressure?	Syringe Example:
What is the relationship between pressure and volume?	Inversely Proportional:



Questions:	Notes:
How can the relationship between pressure and volume be used mathematically?	**Boyle's Law Equation:**
	Example Boyle's Law Problem
	A sample of helium gas in a balloon is compressed from 4.0 L to 2.5 L at a constant temperature. If the pressure of the gas in the 4.0 L volume is 210 kPa, what will the pressure be at 2.5 L?

Summary: