Cornell Notes	Topic/Ol	bjective: Boyle's Law			Name:		
					Class/Period: Chemistry		
AVID® Decades of College Dreams					Date:		
Essential Question: What is the relationship between pressure and volume and how can that							
relationship be used mathematically?							
Questions:		Notes:					
What happens on		Syringe Example:					
the molecular level							
to cause an							
increase in							
pressure?							
<u> </u>							
What is the		Inversely Proportional:					
relationship		Miner sery i i opor tronon					
between pres	sure						
and volume?	-		100	10 —	Gas in a Syringe		
•••			=	-			
			80	0			
			(lb/in ²)	.0			
			 Total Pressure (lb/in ²) 59	-			
			tal Pre	0			
			_ ρ _ 20				
				-			
			_ (0	20 40 60 80 100 120		
			_		Volume (mL)		

Questions:	Notes:
How can the	Boyle's Law Equation:
relationship	
between pressure	
and volume be used	
mathematically?	
	Example Boyle's Law Problem
	A sample of helium gas in a balloon is compressed from 4.0 L to
	2.5 L at a constant temperature. If the pressure of the gas in the
	4.0 L volume is 210 kPa, what will the pressure be at 2.5 L?
Summary:	
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