| Cornell Notes | Topic/Ol | ojective: Activity Series Na | ame: | |
|--|----------|---|---|--|
| | | CI | Class/Period: Chemistry | |
| AVID® Decades of College Dreams | | Da | ate: | |
| Essential Questions: Which metal atoms are most easily oxidized? | | | | |
| | | | | |
| Questions: | | Notes: | Beaker I: Copper in silver nitrate solution | |
| What happens | | | 47 | |
| during a reaction | | | Coppe | |
| between a solid | | | Cu ²⁺ | |
| metal and a metal | | | Cu ²⁺ (Ag/Ag) C, 12+ | |
| ion? | | | Cu ²⁺ Cu ²⁺ | |
| | | | Silver: reduced | |
| | | | Copper: oxidized | |
| | | | Beaker 3: Silver in zinc | |
| Why do reactions | | | iiver | |
| not occur between | | | | |
| a solid metal and a | | | Zinc nitrate | |
| metal ion? | | | | |
| Цом одн мон | | | No reaction | |
| How can you determine if a | | | | |
| metal will rea | | | | |
| WETAI WIII I EA | ivi i | | | |
| | | Use the provided equations to determine where aluminum falls in | | |
| | | the order of activity so far with zinc, copper, and silver. | | |
| | | $Al_{(s)} + Zn(NO_3)_{2 (aq)} \rightarrow No \ neaction$ | | |
| | | $Al(s) + CuNO_{3(aq)} \rightarrow Cu(s) + Al(NO_{3})_{3(aq)}$ | | |
| | | | Zinc More active | |
| | | gives | up electrons to | |
| | | | Copper | |
| | | which give | ves up electrons to | |
| | | | Silver Less active | |

| | Activ | Activity Series | |
|----------|-----------|-----------------|--|
| | Potassium | More active | |
| | Barium | Easily oxidized | |
| | Calcium | - - | |
| | | _ | |
| | Sodium | - | |
| | Magnesium | - | |
| | Aluminum | - | |
| | Zinc | - - | |
| | Chromium | - | |
| | Iron | - - | |
| | Nickel | - | |
| | Copper | - - | |
| | Silver | - | |
| | Mercury | Less active | |
| | Gold | Easily reduced | |
| Summary: | | | |
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