| Cornell Notes | Topic/Objective: Accuracy, Precision, and | | | | Name: | | | |
|--|---|---------------------|------------|--|--------------------------------|-----|---|--|
| | Percent Error | | | | Class/Period: <i>Chemistry</i> | | | |
| Decades of College Dreams | | | | | Date: | | | |
| Essential Question: How do we evaluate the reliability of a measurement? | | | | | | | | |
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| Questions: | | Notes: | | | | | | |
| What is the | | Ассигасу: | | | | | | |
| difference | | | | | | | | |
| between accuracy | | | | | | | | |
| and precision? | | Precision: | | | | | | |
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| HOW IS THE Accept | | Iea Valhe: | | | | | | |
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| mathematic | alluz | enon. | Enor: | | | | | |
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| | Percent Error: |
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| | |
| | $Percent \ Error = \frac{ error }{accepted \ value} \times 100\%$ |
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| | Practice: |
| | Líttle Johnny measured 32.71 g of sodíum chloríde (NaCl), but he should have |
| | had 41.00 g. What was his percent error? |
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| | What is the percent error of a calculated value being 114 g/mL and a reference |
| | book stated the value as 109.5 g/mL? |
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| Summary: | |
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